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# Ce4501 Environmental Engineering Chemical Processes Fall

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Chemical Processes  
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Processes Preventing  
Flint - Environmental  
Engineering: Crash  
Course Engineering  
#29 Environment  
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Engineering What is  
Chemical Engineering?  
Chemical Parameters  
of Water | Lecture 9 |  
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Introduction to  
Chemical Engineering |  
Lecture 1 What's it  
really like to study  
Chemical and*

*Environmental  
Engineering at UQ?  
Biological Parameters  
of Water | Lecture 10 |  
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Treatment of Water |  
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Chemical Parameters |  
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SSC Flocculation |  
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Environmental  
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is Chemical  
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**ENVIRONMENTAL SCIENCE AND ENGINEERING- GATE 2021 COMPLETE GUIDANCE II Aniruddha Sir ( IIT BOMBAY)**

Disinfection of Water | Purification of Water | Part - 04 | Environmental Engineering

Quality of Water | Chemical Parameters | Lecture 7 | Environmental Engineering *Chemical*

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1. Problem 15.3 in the text (p. 334). Using the formation constants in section 15.5.2, find the concentrations of the hydroxo complexes of Cu (including the dimer) in a water at pH 6.0 with total soluble

copper equal to  
 1.0x10<sup>-6</sup>M  
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 Processes Environ.  
 Engineering Course  
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 4505 CE4506  
 Regs,P2,3 Water  
 Resources GE3850  
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 Drinking&WasteWtr  
 CE4507 Distribution. ce  
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 Limnology FW4220  
 Wetlands CE4620 Open  
 Channel Flow CE4630  
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 Environmental  
 Engineering Chemical  
 Processes Problem Set

5 - SOLUTION Fall 2008  
 Due: Monday, 11/10 by  
 5 p.m. 1. Problem 16.1  
 in the text (p. 379) A.  
 Chlorine exists in  
 seven oxidation states.  
 Determine the  
 oxidation state of each  
 compound and order  
 them from the most  
 negative oxidation  
 state to the most  
 positive oxidation  
 state.CE4501  
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 Engineering Chemical  
 Processes ...Ce4501  
 Environmental  
 Engineering Chemical  
 Processes 1 Answer.  
 However, as quantified  
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 Engineering Chemical  
 Processes, Cu will form  
 species such as  
 CuOH<sup>+</sup>, which is the  
 reason the solution will  
 be acidic overall. This  
 could be approximated  
 by thinking of 2 6 as  
 having a pKa of ~6.3.

Because 4 is a stronger acid than  $\text{CuOH}^+$  is a base, the solution is  
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Environmental  
Engineering Chemical  
Processes Problem Set  
1 - SOLUTIONS Fall  
2008 1 . Problem 2.7 in  
the text (p. 31)  
Calculate the molarity  
and normality of the  
following solutions: a)  
20 g/L NaOH MW =  
 $1 \times 23 + 1 \times 16 + 1 \times 1 =$   
40 20 0.5 40 g mole  
mole Molarity Lg L = · =  
0.5 1 0.5 0.5 mole eq  
eq Normality N Lmole L  
= · = = b) 25 g/L  
H<sub>2</sub>SO<sub>4</sub>CE4501  
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Engineering Chemical  
Processes ...Chemical  
Processes CE4501  
Environmental  
Engineering Chemical  
Processes, Fall 2008  
Problem Set 7  
SOLUTIONS 1. The  
accident at the nuclear  
power plant in  
Chernobyl released a  
number of  
radioisotopes to the  
environment. All of

these isotopes  
ultimately ended up in  
soils and sediments.  
Two of the isotopes  
released were 134 Cs  
and 129 I.Ce4501  
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completeness, and one  
problem (selected at  
random) will be graded  
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set contributes 2.5%  
towards your final  
grade. 1. Problem 15.3  
in the text (p. 334).  
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*Ce4501 Environmental  
 Engineering Chemical  
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 Because 4 is a stronger  
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 base, the solution is  
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 $\text{mole}$  Molarity  $\text{Lg L} = \cdot =$   
 $0.5$   $1$   $0.5$   $0.5 \text{ mole eq}$   
 $\text{eq}$  Normality  $\text{N Lmole L}$   
 $= \cdot = =$  b)  $25 \text{ g/L H}_2\text{SO}_4$   
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Environmental  
Research at UNSW  
Chemical Engineering  
Chemical  
Characteristics of  
Water | Environmental  
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