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two orbits it transitions. Energy levels Bohr's atomic model failed to explain the _____ of elements other than hydrogen. Chemistry Chap 5.2 Study Guide (Quantum Theory and the Atom) quantum theory Quantum theory is the theoretical basis of modern physics that explains the nature and behavior of matter and energy on the atomic and subatomic level. The nature and behavior of matter and energy at that level is sometimes referred to as quantum physics and quantum mechanics. What is quantum theory? - Definition from WhatIs.com Quantum mechanics (QM ; also known as quantum physics, quantum theory, the wave mechanical model, or matrix mechanics), including quantum field theory, is a fundamental theory in physics which describes nature at the smallest - including atomic and subatomic - scales. Quantum mechanics - Wikipedia an electron's position cannot be known precisely According to the quantum theory of an atom, in an orbital a. an electron's position cannot be known precisely. b. an electron has no energy. quantum theory Flashcards | Quizlet In chemistry and physics, atomic theory is a scientific theory of the nature of matter, which states that matter is composed of discrete units called atoms. It began as a philosophical concept in ancient Greece and entered the scientific mainstream in the early 19th century when discoveries in the field of chemistry showed that matter did indeed behave as if it were made up of atoms. Atomic theory - Wikipedia The development of quantum theory was arguably many centuries in the making. As early as the 5th Century B.C., the Greek philosophers Democritus and Leucippus first put forward the idea that everything around us was made of tiny

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Quantum mechanics - Wikipedia

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What is quantum theory? - Definition from WhatIs.com

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Chapter 6 Quantum Theory and the Electronic Structure of Atoms

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