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# Electron Probability Lab Answers

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Discussion  
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Atomic Orbital

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Clarification  
on Electron  
Probability Lab  
8.1a  
*Interpreting  
the radial  
probability  
distribution of  
an orbital*

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interpretatio  
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mechanics  
visualized  
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 $N=3$ ,  
 $L=1$ ,  $M=1$ ,  
 $M=1/2$ . This electron is found within the 2p orbital, which has the dumbbell shape, primarily based on the z axis. This electron would share this orbital with

another electron, but would have a negative spin.  
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predicted. The region of space in which the electron can probably be found is called an orbital. Orbitals are often referred to as “electron clouds” because they are not absolute. In fact, there are times when an electron may be found outside the orbital.

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 $M=1$ ,  $M=1/2$ .  
This electron  
is found within  
the 2p orbital,  
which has the  
dumbbell  
shape,  
primarily  
based on the z  
axis. This  
electron would  
share this  
orbital with  
another  
electron, but  
would have a  
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Instead, the quantum-mechanical model for the atom gives

the probabilities of finding an electron in a particular region around the nucleus. In this experiment, we model the probable locations of electrons around the nucleus. We will use dried beans or peas to represent the electrons to help us visualize the regions of low and high density.

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Clarification on Electron Probability Lab 8.1a

*Interpreting the radial probability distribution of an orbital*

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